

Getting To Know The Elements Answer Key

Name _____ KEY _____ Period _____

Get to Know the Elements
Use the periodic table in pictures at <http://elements.alonk.com> to answer the following questions!

How many elements are gaseous at room temperature? 11

Where do the gaseous elements tend to be located on the table?
Toward the right side

What are the top eight elements found in the Earth's crust?
1. oxygen
2. silicon
3. aluminum
4. iron
5. calcium
6. potassium
7. sodium
8. magnesium

What element is used in batteries?
lithium

What element is used in making paint?
cadmium

What is a major element found in our bones?
phosphorus

Which elements are magnetic?
1. iron
2. nickel
3. cobalt

What are the top ten most abundant elements in the human body?
1. hydrogen
2. carbon
3. nitrogen
4. oxygen
5. sodium
6. phosphorus
7. sulfur
8. chlorine
9. potassium
10. calcium

Where do the elements in the human body tend to be located on the table?
Toward the top

What element is used in bicycles?
scandium

What is a major element found in chlorophyll?
magnesium

Which elements are liquid at room temperature?
mercury & bromine

What radioactive element has the lowest atomic number?
Technetium (43)

How many of the elements are radioactive?
37 (including superheavies)

Where do the radioactive elements tend to be located on the table?
Toward the bottom/actinides

Which elements are considered "Noble Metals"?
1. ruthenium
2. rhodium
3. palladium
4. silver
5. osmium
6. iridium
7. platinum
8. gold

Where do the Noble Metals tend to be located on the table?
In the center

Getting to Know the Elements Answer Key: Your Comprehensive Guide

Are you struggling to master the periodic table? Feeling overwhelmed by the sheer number of elements and their properties? You're not alone! Understanding the elements is crucial for success in chemistry and related scientific fields. This comprehensive guide provides a detailed "getting to know the elements answer key," walking you through key concepts, clarifying common misconceptions, and offering resources to help you conquer this fundamental aspect of science. We'll go beyond just memorization, focusing on understanding the underlying principles that govern the behavior of elements.

H2: Unlocking the Periodic Table: Understanding its Structure

The periodic table isn't just a random arrangement of elements; it's a meticulously organized system reflecting the properties and relationships between them. Understanding its structure is the first step to mastering the elements.

Periods (Rows): Elements in the same period have the same number of electron shells. As you move across a period, the atomic number (number of protons) increases, and the properties of the elements change systematically.

Groups (Columns): Elements in the same group have the same number of valence electrons (electrons in the outermost shell). This shared characteristic leads to similar chemical properties. For example, Group 18 (noble gases) are all remarkably unreactive.

Metals, Nonmetals, and Metalloids: The periodic table also visually separates elements into these three broad categories, based on their physical and chemical characteristics. Metals are typically shiny, conductive, and malleable, while nonmetals tend to be brittle and poor conductors. Metalloids exhibit properties of both metals and nonmetals.

H2: Key Properties to Focus On: Beyond Atomic Number

While the atomic number is fundamental, focusing solely on it won't provide a complete understanding. Several other crucial properties define an element's behavior:

Atomic Mass: This represents the average mass of an element's isotopes (atoms with the same number of protons but different numbers of neutrons).

Electronegativity: This describes an atom's ability to attract electrons in a chemical bond. Highly electronegative elements tend to gain electrons, while those with low electronegativity tend to lose them.

Ionization Energy: This is the energy required to remove an electron from an atom. Elements with high ionization energy hold onto their electrons tightly.

Reactivity: This describes how readily an element participates in chemical reactions. Highly reactive elements readily form bonds with other elements.

H3: Mastering Element Families: Groups and Their Characteristics

Understanding element families simplifies the learning process. Focusing on the characteristic traits of each group allows for better prediction of an element's behavior.

Alkali Metals (Group 1): Highly reactive, readily lose one electron to form +1 ions.

Alkaline Earth Metals (Group 2): Reactive, but less so than alkali metals; lose two electrons to form +2 ions.

Halogens (Group 17): Highly reactive nonmetals; readily gain one electron to form -1 ions.

Noble Gases (Group 18): Extremely unreactive due to their full valence electron shells.

H2: Utilizing Resources Effectively: Going Beyond the Textbook

Textbooks offer a foundational understanding, but several additional resources can significantly enhance your learning:

Interactive Periodic Tables: Many online interactive periodic tables provide detailed information about each element, including images, videos, and links to further research.

Educational Videos: YouTube and other platforms offer numerous educational videos explaining the periodic table and the properties of individual elements.

Online Quizzes and Practice Tests: Regular practice is key to solidifying your understanding. Many websites offer quizzes and tests to reinforce your learning.

Study Groups: Collaborating with peers can provide valuable insights and help clarify any confusion.

H2: Common Misconceptions and How to Avoid Them

Several common misconceptions surrounding the elements can hinder understanding. Addressing these misconceptions head-on is crucial for effective learning:

Atomic Number vs. Atomic Mass: Many students confuse these two concepts. Remember, atomic number is the number of protons, while atomic mass is the average mass of isotopes.

Reactivity and Electronegativity: These are related but distinct concepts. Reactivity is about the tendency to react, while electronegativity is about attracting electrons within a bond.

Metals vs. Nonmetals: Understanding the defining characteristics of each category is key to accurately classifying elements.

H2: Putting it all Together: A Practical Approach

The key to mastering the elements is a combination of understanding fundamental principles, utilizing effective resources, and consistent practice. Don't just memorize; strive to understand the underlying reasons behind the properties and behaviors of each element. By combining conceptual learning with practical application, you'll build a solid foundation in chemistry and related scientific fields.

Conclusion

Mastering the elements requires dedication and a systematic approach. By understanding the structure of the periodic table, focusing on key properties, utilizing available resources, and actively addressing common misconceptions, you can unlock a deeper understanding of this fundamental aspect of chemistry. Remember, the journey to mastering the elements is a process of continuous learning and refinement.

FAQs

1. Where can I find a reliable online periodic table? Many websites offer interactive periodic tables, including Ptable.com and Royal Society of Chemistry's periodic table. Choose one with detailed information and interactive features.
2. How can I memorize the elements effectively? Instead of rote memorization, focus on understanding the periodic trends and the properties of element families. Use flashcards or mnemonic devices to aid your memory, but prioritize understanding over memorization.
3. What are some good resources for visual learners? Videos on YouTube and interactive periodic tables are excellent resources for visual learners. Look for videos that use animations and diagrams to explain complex concepts.
4. How can I improve my problem-solving skills related to the elements? Practice solving problems that involve predicting chemical reactions, calculating atomic masses, and identifying elements based on their properties. Use online quizzes and textbooks to find practice problems.
5. Are there any mobile apps that can help me learn about the elements? Yes, many mobile apps offer interactive periodic tables and quizzes. Search your app store for "periodic table" or "chemistry elements" to find suitable options.

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Glenda Lappan, James Taylor Fey, William Morley Fitzgerald, Susan N. Friel, Elizabeth Difanis Phillips, 2004

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from Concordia University, Canada; Prof N A Strelkov and V L Dol'nikov from Yaroslavl State University, Russia; Prof Chin-Chen Chang and Ching-Yun Chang from Taiwan; Prof S S Pandey from R D University, India; and Prof I L Bloshanskii from Moscow State Regional University, Russia.

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easy-to-use workbook style, the Field Guide provides nearly 100 skill-building exercises of varying types to help you uncover what measurement work your organization is doing; assess organizational maturity and gaps; understand how to apply specific concepts; and determine what's right for your organization moving forward. It also offers interview questions to better understand users' wants and needs; case study exercises to test your knowledge gaps; and reflection questions that focus on your deepening knowledge and skill. You can write your answers in the book or use the resource on ATD's website. Achieving measurement maturity is a change effort requiring commitment and discipline. Understanding your current capabilities and gaps is an essential first step followed by determining where your organization wants to go in this area. Once both are achieved, you will be able to develop your desired measurement and reporting strategy—the bridge between where you are today and where you aspire to be. The Measurement Demystified Field Guide is that bridge.

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