

Lewis Structure Worksheet With Answers

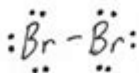
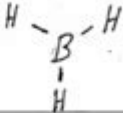
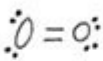
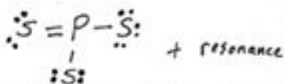
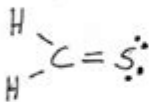
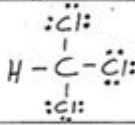
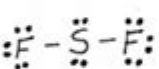
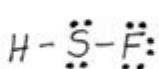
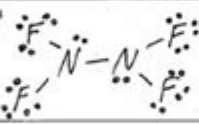
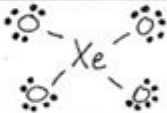
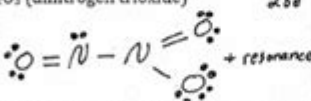
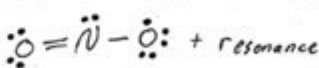
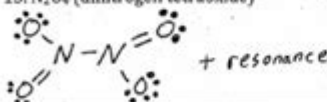
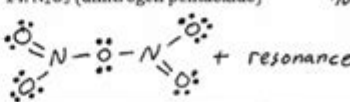
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Lewis Structure Worksheet #2

Key

Draw Lewis Structures for each of the following molecules or ions:

1. Br ₂ 14 	2. BH ₃ 6 
3. O ₂ 12 	4. PS ₃ ⁻ 24 
5. H ₂ CS 12 	6. CHCl ₃ 26 
7. SF ₂ 20 	8. SHF 14 
9. N ₂ F ₄ 38 	10. XeO ₄ 32 
11. N ₂ O ₃ (dinitrogen trioxide)* 28e 	12. NO ₂ (nitrogen dioxide)* 17e 
13. N ₂ O ₄ (dinitrogen tetroxide)* 34 	14. N ₂ O ₅ (dinitrogen pentaoxide)* 40e 

*Problems 11-14 are more challenging and may have more than one correct structure. Please note that if a molecule has an odd number of valence electrons, at least one atom in the molecule will NOT fulfill the octet rule.

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Lewis Structure Worksheet with Answers: Mastering Chemical Bonding

Are you struggling to grasp the concept of Lewis structures? Do those dots and lines representing valence electrons and bonds leave you feeling confused? You're not alone! Many students find Lewis structures challenging, but mastering them is crucial for understanding chemical bonding and predicting molecular geometry. This comprehensive guide provides a Lewis structure worksheet with answers, along with detailed explanations to help you conquer this important chemistry concept. We'll walk you through the steps, provide examples, and offer resources to solidify your

understanding. Let's dive in!

Understanding Lewis Structures: A Quick Recap

Before we tackle the worksheet, let's briefly review the fundamentals of Lewis structures. Lewis structures, also known as Lewis dot diagrams, are visual representations of the valence electrons in an atom or molecule. They help us understand how atoms share electrons to form covalent bonds and predict the overall shape of a molecule. Key components include:

Valence Electrons: These are the electrons in the outermost shell of an atom, which participate in bonding.

Octet Rule: Most atoms strive to achieve a stable electron configuration with eight valence electrons (like a noble gas). Exceptions exist, particularly for elements like hydrogen and boron.

Lone Pairs: These are pairs of valence electrons that are not involved in bonding.

Bonding Pairs: These are pairs of valence electrons shared between atoms to form a covalent bond.

Steps to Draw a Lewis Structure

Drawing a Lewis structure involves several key steps:

1. **Count Valence Electrons:** Determine the total number of valence electrons for all atoms in the molecule or ion.
2. **Identify the Central Atom:** Usually, the least electronegative atom (excluding hydrogen) is the central atom.
3. **Arrange Atoms:** Place the central atom in the center and arrange the other atoms around it.
4. **Connect Atoms with Single Bonds:** Draw single bonds (one shared electron pair) between the central atom and surrounding atoms.
5. **Distribute Remaining Electrons:** Place the remaining valence electrons as lone pairs on the surrounding atoms to satisfy the octet rule (or duet rule for hydrogen).
6. **Satisfy the Octet Rule for the Central Atom:** If the central atom does not have an octet, form double or triple bonds by moving lone pairs from surrounding atoms to form additional shared pairs.
7. **Check Formal Charges (Optional):** Calculate formal charges to determine the most stable Lewis structure.

Lewis Structure Worksheet with Answers: Practice Problems

Now, let's put our knowledge into practice with a Lewis structure worksheet. Remember to follow

the steps outlined above for each problem.

Worksheet Problems:

1. Draw the Lewis structure for H_2O .
2. Draw the Lewis structure for CO_2 .
3. Draw the Lewis structure for NH_3 .
4. Draw the Lewis structure for CH_4 .
5. Draw the Lewis structure for O_2 .
6. Draw the Lewis structure for SO_4^{2-} (sulfate ion).
7. Draw the Lewis structure for NO_3^- (nitrate ion).
8. Draw the Lewis structure for PCl_5 .
9. Draw the Lewis structure for SF_6 .
10. Draw the Lewis structure for HCN .

Worksheet Answers and Explanations:

(Detailed step-by-step solutions for each problem would be provided here. Due to the length constraint, these are omitted. A downloadable PDF with complete solutions could be linked here.)

Beyond the Basics: Resonance and Exceptions to the Octet Rule

While the worksheet covers fundamental Lewis structures, it's crucial to understand more advanced concepts like resonance and exceptions to the octet rule.

Resonance Structures:

Some molecules have multiple valid Lewis structures that differ only in the placement of electrons. These are called resonance structures, and the actual molecule is a hybrid of these structures.

Exceptions to the Octet Rule:

Some molecules have central atoms with fewer than eight (e.g., boron in BF_3) or more than eight (e.g., phosphorus in PCl_5) valence electrons. These exceptions are often explained by considering factors beyond the simple octet rule.

Conclusion

Mastering Lewis structures is fundamental to understanding chemistry. This Lewis structure worksheet with answers provides a valuable tool for practicing and improving your skills. Remember to review the steps carefully, and don't hesitate to utilize online resources and your textbook for further clarification. Consistent practice will build your confidence and help you accurately predict molecular properties.

FAQs

Q1: What is the difference between a single, double, and triple bond in a Lewis structure?

A1: A single bond represents one shared electron pair, a double bond represents two shared electron pairs, and a triple bond represents three shared electron pairs between two atoms.

Q2: How do I determine the central atom in a molecule?

A2: Generally, the least electronegative atom (excluding hydrogen) acts as the central atom. However, there can be exceptions, especially with larger molecules.

Q3: What are formal charges, and why are they important?

A3: Formal charges are a way to assess the distribution of electrons in a Lewis structure. They help determine the most stable structure by minimizing formal charges.

Q4: What resources are available to help me learn more about Lewis structures?

A4: Many excellent online resources, including educational videos, interactive simulations, and practice problems, are available. Your chemistry textbook and instructor are also valuable resources.

Q5: How can I tell if I've drawn the correct Lewis structure?

A5: Check that all atoms (except hydrogen) have a full octet (or duet for hydrogen), the total number of valence electrons is correct, and the formal charges are minimized. Comparing your structure to established resources can also be helpful.

William R. Robinson, Klaus Hellmut Theopold, 2019-02-14 Chemistry 2e is designed to meet the scope and sequence requirements of the two-semester general chemistry course. The textbook provides an important opportunity for students to learn the core concepts of chemistry and understand how those concepts apply to their lives and the world around them. The book also includes a number of innovative features, including interactive exercises and real-world applications, designed to enhance student learning. The second edition has been revised to incorporate clearer, more current, and more dynamic explanations, while maintaining the same organization as the first edition. Substantial improvements have been made in the figures, illustrations, and example exercises that support the text narrative. Changes made in Chemistry 2e are described in the preface to help instructors transition to the second edition.

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