

Molarity Worksheet Answer Key

Molarity Worksheet #1

name _____

1. What does molarity mean?

$$\frac{\text{Number of moles of solute}}{\text{1 liter solution}}$$

2. What is the molarity of a solution that contains 4.53 moles of lithium nitrate in 2.85 liters of solution?

$$\frac{4.53 \text{ mol LiNO}_3}{2.85 \text{ L soln}} = 1.59 \text{ M LiNO}_3$$

3. What is the molarity of a solution that contains 0.00372 moles hydrochloric acid in 2.39×10^{-2} liters of solution?

$$\frac{0.00372 \text{ mol HCL}}{2.39 \times 10^{-2} \text{ L soln}} = 0.156 \text{ M HCL}$$

4. A flask contains 85.5 g $\text{C}_{12}\text{H}_{22}\text{O}_{11}$ (sucrose) in 1.00 liters of solution. What is the molarity?

$$\frac{85.5 \text{ g sucrose} \times 1 \text{ mol sucrose}}{1.00 \text{ L soln} \quad 342.34 \text{ g sucrose}} = 0.250 \text{ M sucrose}$$

5. A beaker contains 214.2 grams osmium (III) fluoride in 0.0673 liters of solution. What is the molarity?

$$\frac{214.2 \text{ g OsF}_3 \times 1 \text{ mol OsF}_3}{0.0673 \text{ L soln} \quad 247.23 \text{ g OsF}_3} = 12.9 \text{ M OsF}_3$$

6. Calculate the molarity if a flask contains 1.54 moles potassium sulfate in 125 ml of solution.

$$\frac{1.54 \text{ mol K}_2\text{SO}_4}{0.125 \text{ L soln}} = 12.3 \text{ M K}_2\text{SO}_4$$

Molarity Worksheet Answer Key: Mastering Solution Concentration Calculations

Are you struggling to understand molarity calculations? Feeling overwhelmed by those tricky molarity worksheets? You're not alone! Many students find molarity – a crucial concept in chemistry – challenging. This comprehensive guide provides you with not only a detailed explanation of molarity but also serves as a de facto molarity worksheet answer key, helping you check your work and solidify your understanding. We'll walk you through various molarity problems, offering step-by-step solutions and tips to conquer even the most complex calculations. Let's dive in!

Understanding Molarity: A Quick Recap

Before we jump into the molarity worksheet answer key examples, let's briefly review the concept of molarity. Molarity (M) is a measure of the concentration of a solution. It's defined as the number of moles of solute (the substance being dissolved) per liter of solution. The formula is:

Molarity (M) = moles of solute / liters of solution

Common Molarity Problems and Their Solutions (Molarity Worksheet Answer Key Examples)

This section acts as your comprehensive molarity worksheet answer key, offering solutions to common problem types.

1. Calculating Molarity Given Moles and Volume

Problem: Calculate the molarity of a solution containing 0.5 moles of sodium chloride (NaCl) dissolved in 250 mL of water.

Solution:

1. Convert volume to liters: $250 \text{ mL} (1 \text{ L} / 1000 \text{ mL}) = 0.25 \text{ L}$
2. Apply the molarity formula: $M = 0.5 \text{ moles} / 0.25 \text{ L} = 2.0 \text{ M}$

Answer: The molarity of the solution is 2.0 M.

2. Calculating Moles Given Molarity and Volume

Problem: How many moles of potassium hydroxide (KOH) are present in 500 mL of a 0.2 M solution?

Solution:

1. Convert volume to liters: $500 \text{ mL} (1 \text{ L} / 1000 \text{ mL}) = 0.5 \text{ L}$
2. Rearrange the molarity formula to solve for moles: $\text{moles} = \text{Molarity} \times \text{liters}$
3. Calculate moles: $\text{moles} = 0.2 \text{ M} \times 0.5 \text{ L} = 0.1 \text{ moles}$

Answer: There are 0.1 moles of KOH in the solution.

3. Calculating Volume Given Molarity and Moles

Problem: What volume of a 1.5 M solution of sulfuric acid (H_2SO_4) contains 0.75 moles of H_2SO_4 ?

Solution:

1. Rearrange the molarity formula to solve for volume: $\text{liters} = \text{moles} / \text{Molarity}$
2. Calculate volume: $\text{liters} = 0.75 \text{ moles} / 1.5 \text{ M} = 0.5 \text{ L}$
3. Convert liters to milliliters (optional): $0.5 \text{ L} (1000 \text{ mL} / 1 \text{ L}) = 500 \text{ mL}$

Answer: 500 mL of the 1.5 M H_2SO_4 solution contains 0.75 moles of H_2SO_4 .

4. Dilution Problems: $M_1V_1 = M_2V_2$

Dilution involves decreasing the concentration of a solution by adding more solvent. The equation $M_1V_1 = M_2V_2$ is used, where:

M_1 = initial molarity

V_1 = initial volume

M_2 = final molarity

V_2 = final volume

Problem: 100 mL of a 3.0 M solution of HCl is diluted to a final volume of 500 mL. What is the final molarity?

Solution:

1. Apply the dilution formula: $(3.0 \text{ M})(100 \text{ mL}) = M_2(500 \text{ mL})$
2. Solve for M_2 : $M_2 = (3.0 \text{ M } 100 \text{ mL}) / 500 \text{ mL} = 0.6 \text{ M}$

Answer: The final molarity of the HCl solution is 0.6 M.

Tips for Success with Molarity Calculations

Units are crucial: Always pay close attention to units (moles, liters, milliliters).

Organize your work: Write down the given information, the formula, and your calculations neatly.

Check your answers: Make sure your answer makes sense in the context of the problem.

Practice regularly: The more you practice, the more comfortable you'll become with molarity calculations.

Conclusion

Mastering molarity is essential for success in chemistry. By understanding the fundamental concepts and practicing with various problem types, you can confidently tackle any molarity worksheet. This guide, acting as your personalized molarity worksheet answer key, provides a solid foundation and solutions to common challenges. Remember to practice consistently, and you'll soon be an expert in molarity calculations!

Frequently Asked Questions (FAQs)

1. What is the difference between molarity and molality? Molarity is moles of solute per liter of solution, while molality is moles of solute per kilogram of solvent.
2. Can molarity be negative? No, molarity is always a positive value because it represents the concentration of a substance.
3. How do I handle molarity problems with multiple solutes? You need to calculate the molarity of each solute individually. The total molarity is not simply the sum of individual molarities unless the solutes are fully independent (no interactions).
4. What if my answer doesn't exactly match the answer key? Small discrepancies can arise due to rounding errors in calculations. As long as your answer is reasonably close (within a few percentage points), it's likely correct.
5. Where can I find more molarity practice problems? Many chemistry textbooks and online resources offer additional molarity practice problems and worksheets. You can also search for "molarity practice problems with answers" on the internet.

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25 MIN CARDIO HIIT para perder barriga rápido, sem ... - YouTube

Está procurando um cardio hiit para perder a barriga que você possa fazer em casa e sem equipamentos? Este treino de 25 minutos é perfeito para você!

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