

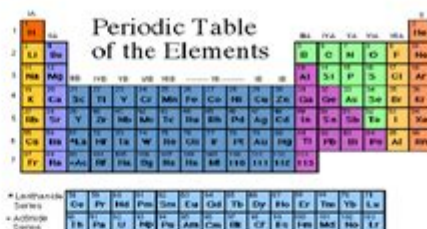
Periodic Table Webquest Answer Key

Periodic Table WebQuest

Name: _____

Date: _____

Pd: _____

A standard periodic table of elements, color-coded by groups. It includes element symbols, atomic numbers, and names. The title "Periodic Table of the Elements" is centered at the top.

Instructions: In each of the websites listed, find the information asked for.

Part I: **"Major Players"** Go to

<http://allperiodictables.com/ClientPages/AAEpages/aacHistory.html>

For this site, write the contributions of each person(s) toward the development of the Periodic Table of Elements.

Scientist	Contribution to the development of the periodic table
Greek thinkers	
Lavoisier	
John Dalton	
Doberiner	
De Chancourtois	
Cannizzaro	
Newlands	
Meyer	
Mendeleev	
Mosely	
Seaborg	

Part II: **"Get Organized Periodically"** Go to

http://www.chem4kids.com/files/elem_pertable.html

1. Why are the elements placed in specific places on the Periodic Table? _____

1

Periodic Table WebQuest Answer Key: A Comprehensive Guide

Are you struggling to complete your periodic table webquest? Feeling overwhelmed by the sheer amount of information packed into this essential chemistry tool? Don't worry, you're not alone! Many students find navigating the complexities of the periodic table challenging. This comprehensive guide provides a structured approach to answering common periodic table webquest questions, eliminating the frustration and helping you achieve a top grade. We'll delve into key concepts, providing insightful answers and explaining the reasoning behind them. This isn't just an answer key; it's a learning tool designed to enhance your understanding of the periodic table.

Understanding Your WebQuest: A First Step

Before diving into specific answers, it's crucial to understand the nature of your webquest. Different assignments have varying levels of detail and specific focus areas. Carefully review your instructor's instructions and the specific questions posed in your webquest. This initial step will guide your research and ensure you address all the necessary elements. Look for keywords within the questions themselves – these keywords will often point you towards the right information online.

Navigating the Periodic Table: Key Concepts and Terminology

The periodic table organizes chemical elements based on their atomic number, electron configuration, and recurring chemical properties. Understanding these key concepts is vital to answering most webquest questions.

1. Atomic Number and Atomic Mass:

Atomic Number (Z): This represents the number of protons in an atom's nucleus. It uniquely identifies each element. Understanding atomic number is crucial for locating elements on the table.

Atomic Mass (A): This is the average mass of an atom of an element, considering the different isotopes. It's usually a decimal number because it reflects the relative abundance of isotopes.

2. Groups (Columns) and Periods (Rows):

Groups: Elements in the same group (column) share similar chemical properties due to having the same number of valence electrons (electrons in the outermost shell).

Periods: Elements in the same period (row) have the same number of electron shells.

3. Metals, Nonmetals, and Metalloids:

Metals: Generally located on the left side of the periodic table, metals are characterized by their conductivity, malleability, and ductility.

Nonmetals: Typically found on the right side of the periodic table, nonmetals are poor conductors of electricity and heat.

Metalloids: These elements possess properties of both metals and nonmetals and are located along the staircase-like line separating metals and nonmetals.

Common WebQuest Questions and Answers

While the specific questions in your webquest will vary, here are some common themes and example answers to illustrate the approach:

Q: What is the atomic number and symbol of oxygen?

A: The atomic number of oxygen is 8, and its symbol is O.

Q: Describe the properties of alkali metals.

A: Alkali metals (Group 1) are highly reactive metals that readily lose one electron to form a +1 ion. They are soft, silvery-white, and have low densities.

Q: Locate and identify three elements that are halogens.

A: Fluorine (F), Chlorine (Cl), and Bromine (Br) are halogens, belonging to Group 17. They are highly reactive nonmetals.

Q: Explain the trend of electronegativity across a period.

A: Electronegativity generally increases across a period from left to right. This is because the effective nuclear charge increases, attracting valence electrons more strongly.

Q: What is the difference between an atom and an ion?

A: An atom is electrically neutral, with equal numbers of protons and electrons. An ion is an atom or molecule that has gained or lost one or more electrons, resulting in a net positive (cation) or negative (anion) charge.

Remember: This is a guide. Use your own research and the resources provided in your webquest to verify these answers and expand your understanding. Don't just copy and paste; understand the concepts!

Utilizing Online Resources Effectively

To ensure accurate answers, utilize reputable online resources like the Royal Society of Chemistry's website, the National Institute of Standards and Technology (NIST) database, and educational websites from trusted universities. Cross-reference information from multiple sources to ensure accuracy.

Conclusion

Completing a periodic table webquest can be a rewarding experience. By understanding the fundamental concepts of the periodic table and using this guide as a structured framework, you can navigate the assignment confidently and deepen your understanding of chemistry's foundational tool. Remember to always cite your sources and showcase your understanding beyond simply providing answers. Good luck!

FAQs

1. My webquest asks about isotopes. How do I approach that? Isotopes are atoms of the same element with the same number of protons but a different number of neutrons. Focus on the definitions and look for information on the relative abundance of isotopes when calculating atomic mass.
2. I'm struggling with electron configurations. Where can I find help? Many online resources offer interactive tools and tutorials on electron configurations. Search for "electron configuration tutorial" to find helpful visuals and explanations.
3. My webquest includes questions about chemical reactions. How do I approach these? Refer to your chemistry textbook or online resources to understand basic chemical reaction types (e.g., synthesis, decomposition, single displacement, double displacement).
4. What if my webquest has unique questions not covered here? Use the keywords from those questions in your online searches. You'll likely find information on specialized websites or academic journals.
5. How can I check my answers for accuracy? Compare your answers with information from multiple reputable sources. Consult your textbook, teacher, or other reliable online resources to ensure accuracy.

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Instruction That Works answered some vital questions about 21st century teaching and learning: What are the best ways to incorporate technology into the curriculum? What kinds of technology will best support particular learning tasks and objectives? How does a teacher ensure that technology use will enhance instruction rather than distract from it? This revised and updated second edition of that best-selling book provides fresh answers to these critical questions, taking into account the enormous technological advances that have occurred since the first edition was published, including the proliferation of social networks, mobile devices, and web-based multimedia tools. It also builds on the up-to-date research and instructional planning framework featured in the new edition of Classroom Instruction That Works, outlining the most appropriate technology applications and resources for all nine categories of effective instructional strategies: * Setting objectives and providing feedback * Reinforcing effort and providing recognition * Cooperative learning * Cues, questions, and advance organizers * Nonlinguistic representations * Summarizing and note taking * Assigning homework and providing practice * Identifying similarities and differences * Generating and testing hypotheses Each strategy-focused chapter features examples—across grade levels and subject areas, and drawn from real-life lesson plans and projects—of teachers integrating relevant technology in the classroom in ways that are engaging and inspiring to students. The authors also recommend dozens of word processing applications, spreadsheet generators, educational games, data collection tools, and online resources that can help make lessons more fun, more challenging, and—most of all—more effective.

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familiarize the advanced undergraduate and beginning graduate student with basic concepts, theories, models, and experimental techniques for polymer solutions; and to provide a reference for researchers working in the area of polymer solutions as well as those in charge of chromatographic characterization of polymers. The author's incorporation of recent advances in the instrumentation of size-exclusion chromatography, the method by which polymers are analyzed, renders the text particularly topical. Subjects discussed include: Real, ideal, Gaussian, semirigid, and branched polymer chains Polymer solutions and thermodynamics Static light scattering of a polymer solution Dynamic light scattering and diffusion of polymers Dynamics of dilute and semidilute polymer solutions Study questions at the end of each chapter not only provide students with the opportunity to test their understanding, but also introduce topics relevant to polymer solutions not included in the main text. With over 250 geometrical model diagrams, *Polymer Solutions* is a necessary reference for students and for scientists pursuing a broader understanding of polymers.

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periodic table webquest answer key: Forensic Science for High School Barbara Deslich, John Funkhouser, Kendall/Hunt Publishing Company, 2009

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